

From an old drug to a new one: Synthesis of valproate from 5,5-dipropylbarbituric acid

OLDŘICH FARSA, PAVOL BRKA

Department of Chemical Drugs, Faculty of Pharmacy University of Veterinary and Pharmaceutical Sciences, Brno, Czech Republic

Introduction

Barbiturates are old-fashioned and mainly obsolete sedatives, hypnotics and antiepileptics. Only phenobarbital (Phenaemal[®], Phenaemal-tten[®]) remains approved as a reserve medicine for epilepsy. Symmetrical barbiturates such as allobarbital (5,5-diallylbarbituric acid) or 5,5-dipropylbarbituric acid were mainly used in irrational combinations with weak analgesics and are completely obsolete. On the contrary, valproic acid and its salts are modern and popular therapeutics which are used not only as antiepileptics. Symmetrical barbiturates were being synthesized in our department practical courses during the 1990's. We have thus hardly utilizable reserves of these compounds. Some of them are classified as psychotropic compounds in accordance with the Addictive Compounds Act¹⁾ and the Government Regulation Concerning the Lists of Addictive Compounds²⁾. It is desirable to get rid of such reserves in accordance with the law. That is why we have proposed to use our old department reserves of 5,5-dipropylbarbituric acid as the starting material for a new experiment in our Medicinal Chemistry practical courses in which valproic acid will be prepared.

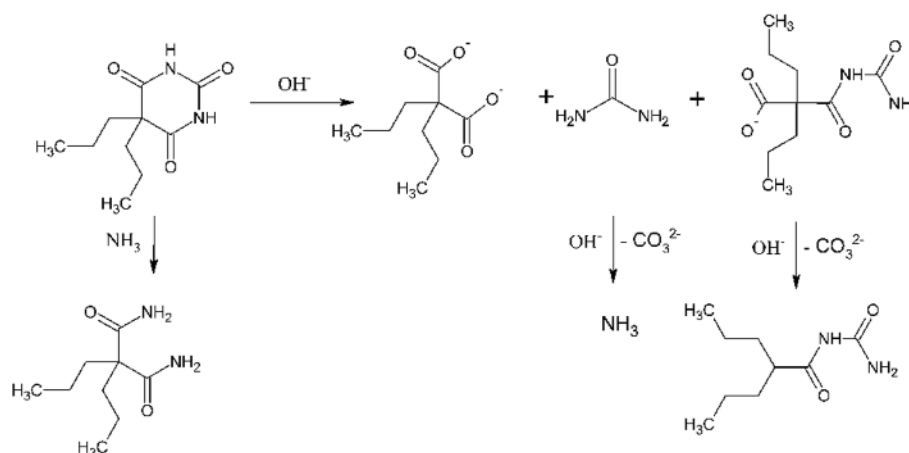


Fig. 1. Hydrolysis and other reactions of 5,5-dipropylbarbituric acid in an alkaline hydroxide solution

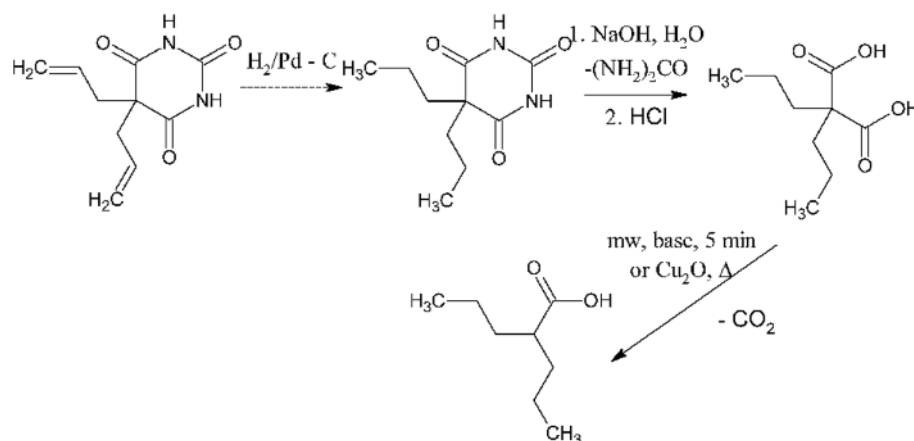


Fig. 2. The overall scheme of synthesis of valproic acid from 5,5-dipropylbarbituric acid (or from allobarbital)

Experimental methods

Experiments involved optimization of two synthetic steps: hydrolysis of 5,5-dipropylbarbituric acid to 2,2-dipropylmalonic acid and decarboxylation of the latter to 2-propylpentanoic, e.g. valproic acid. The first synthetic step had to be optimized because the composition of products of the alkaline hydrolysis depends on the reaction conditions. Such hydrolysis can result in 2,2-dipropylmalonuric acid or in mono- or diamide of 2,2-dipropylmalonic acid except or instead of the desirable 2,2-dipropylmalonic acid^{3,4)} (Fig. 1). These problems were solved by usage of enough excess of sodium hydroxide, its optimal concentration and the satisfactory reaction time. The second synthetic step is

a thermal decomposition – decarboxylation of 2,2-dipropylmalonic acid. It has been optimized for performance in a microwave reactor where it, under catalysis with pyridine, takes only a few minutes⁵). The product is then isolated by dissolving of the reaction mixture in water, acidification and extraction of the product into diethyl ether. In need it can be distilled under reduced pressure, typically using a ball-to-ball distillation apparatus. We had also found conditions for the conventional version of this reaction which requires copper(I) oxide catalysis⁶). This version takes about 30 minutes (Fig. 2).

Results and discussion

The two-stage synthetic procedure for synthesis of valproic acid has been developed. The first stage is alkaline hydrolysis of 5,5-dipropylbarbituric acid to 2,2-dipropylmalonic acid in a larger amount. This step will be performed by a laboratory technician to avoid the contact of students with the barbiturate. The second stage is the thermal decarboxylation of 2,2-dipropylmalonic acid to valproic acid. This step will be performed by students either in a microwave reactor or by a conventional heating. It must be noted that one-pot syntheses of α -substituted alkanic acids directly from a barbiturate have also been reported⁷) but they take at least 24 hours which is inapplicable in practical classes. When the reserve of 5,5-dipropylbarbituric acid will be consumed, we will hydrogenate our reserve allobarbitol to get it (Fig. 2). After consumption of allobarbitol, the first stage will be replaced with a conventional synthesis of 2,2-dipropylmalonic acid from diethyl malonate.

Conclusions

A simple two-step synthetic procedure which enables the synthesis of valproic acid from 5,5-dipropylbarbituric acid suitable for practical courses has been developed. The slower first step, hydrolysis of 5,5-dipropylbarbituric acid to 2,2-dipropylmalonic acid, is intended to be performed by a laboratory technician to prepare a reserve of this intermediate while the second step, decarboxylation of 2,2-dipropylmalonic acid to 2-propylpentanoic (valproic) acid will be made by students.

Conflicts of interest: none.

References

1. Addictive Compounds Act No. 167/1998 Coll.
2. Government Regulation Concerning the lists of Addictive Compounds No. 463/2013 Coll.
3. **Aspelund H., Skoglund L.** The stability of the sodium salts of some soporifics of the barbituric acid series in solution. *Acta Acad. Aboensis Math. et Phys.* 1937; 10, 14–36.
4. **Mokrosz J., Bojarski J.** Hydrolysis of barbituric acid derivatives. Part IV. Hydrolysis of derivatives without a shielding steric effect. *Polish Journal of Chemistry* 1982; 56, 491–500.
5. **Helavi V. B., Solabannavar S. B., Desai U. V., Mane R. B.** Microwave assisted hydrolysis of Meldrum's acid derivatives and decarboxylation of derived malonic acids. *J. Chem. Research (S)*, 2003; 2003, 174–175.
6. **Toussaint O., Capdevielle P., Maumy M.** The Copper(I)-Catalyzed Decarboxylation of Malonic Acids; A New Mild and Quantitative Method. *Synthesis* 1986; 1986, 1029–1031.
7. **Barton H., Mokrosz J., Bojarski J., Klimczak M.** Photochemical degradation of barbituric acid derivatives. Part 1. Products of photolysis and hydrolysis of pentobarbital. *Pharmazie* 1980; 35, 155–158.